













Robotics
Club



LAVA LAMP

Procedure:

Fill the bottle with a quarter full of water. Use the funnel to fill the water so that water doesn't spill outside. (60mL)

Add vegetable oil in the quarter-filled water bottle until it is full (120mL).

Now, wait for a few minutes until the water and oil get separate.

Add food coloring into the bottle.

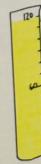
Now we can see that food coloring goes through oil and gets mixed with water.

Cut the fizz tablet into a few smaller pieces and drop only one piece in the bottle. This will start fizzing and appears like a lava lamp.

Whenever required, drop another piece of the tablet to make a few of bubbling in case the lava lamp stops.

Materials:

>Oil $\frac{3}{4}$ cup (20mL)



>food Coloring (3-4 drops)



>Alka-Seltzer (1 tab)



Application to the real world: Our lava lamp project explores various concepts that are used in the real world. Immiscibility is the inability of substances to mix. This is relevant to many science uses. Oil and water do not mix. Emulsifiers are substances that help oil and water mix more effectively. As well as ice particles such as snowflakes and icebergs don't melt if they interact with other materials, including materials like the snowman that don't produce water.

Objective

Explaining how substances with differing polarities interact with each other while showing that science can still be cool & hip!

Fun fact!

Oil molecules are nonpolar. Meanwhile water molecules are electrically polar (+ Hydrogen pole on one side and - negative Oxygen pole on the other). But since oil doesn't have those charges the oil is unaffected by the partial electrical charges of water. Another reason as to why oil and water don't mix.

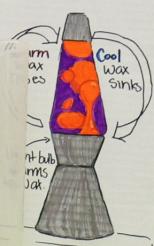
Origins:

- Edward Craven-Walther invented lava lamps in 1963

- Originally intended as decoration, a luxury item but turned into a psychedelic counterculture accessory in the 1960s by hippies.

- In contrast to the experimental lava lamps, actual lava lamps use the heat that comes from a light bulb in the bottom that heat up the wax. This then creates convection currents that makes globs of wax rise as they heat up and later sink as the wax cools down.

Results!



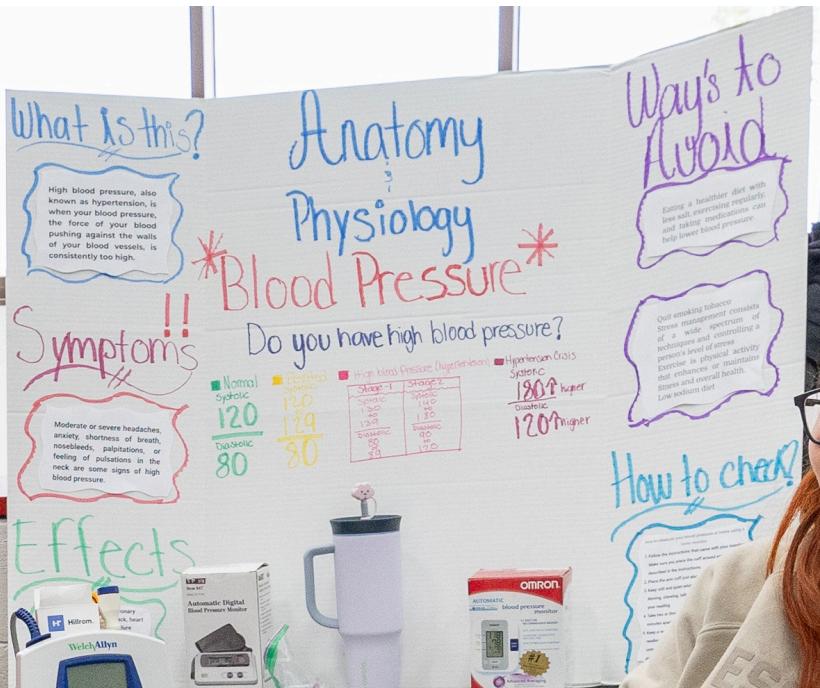
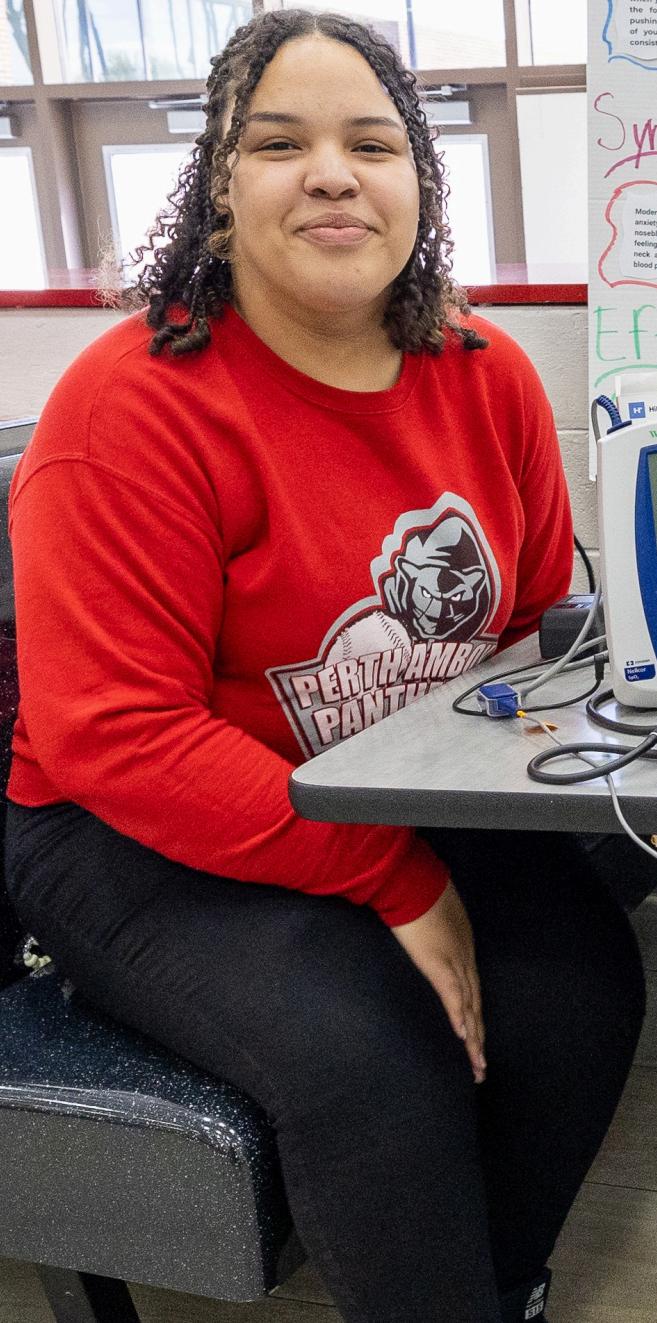














13

Al

Aluminum

26.9815

Alexander Zayas

PER: 3

Trends

1. The Atomic Radius of Aluminum is 1.45 pm. The Atomic Radius is the total distance from the nucleus to the outermost orbital of its electrons.
2. The Ionic Radius of Aluminum is 0.95 pm. The Ionic Radius is the distance from the nucleus to the outermost orbital of an ion, or in which case, the outermost orbital of a cation.
3. This is the density of aluminum at 20°C. The Density of Aluminum is the amount of mass contained in a specific volume of aluminum.
4. The Ionization Energy of Aluminum is 16.02 eV. The Ionization Energy is the amount of energy required to remove an electron from an atom.



19

K

Potassium

39.098



15

P

Phosphorus

30.97